ARTICLE

New carvacrol and thymol derivatives as potential insecticides: synthesis, biological activity, computational studies and nanoencapsulation

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Carolina M. Natal[a], Maria José G. Fernandes[a], Nuno F. S. Pinto[a], Renato B. Pereira[b], Tatiana F. Vieira[c,d], Ana Rita O. Rodrigues[e], David M. Pereira[b], Sérgio F. Sousa[c,d], A. Gil Fortes[a], Elisabete M. S. Castanheira[e], M. Sameiro T. Gonçalves[*a]

New compounds with potential insecticide activity were synthesized by structural modifications performed in the monoterpenoids phenolic moieties of carvacrol and thymol, resulting in a set of derivatives with the ether function containing the propyl, chloropropyl or hydroxypropyl chains, as well as a bicyclic ether with an unsaturated chain containing a carboxylic acid terminal. In addition, an analogue of carvacrol and thymol isomers bearing methoxyl, 1-hydroxyethyl and (3-chlorobenzoyl)oxy, instead of the three original methyl groups, was also synthesized. Several structural changes that resulted in diminished insecticide activity have been identified, but two significantly active molecules have been synthesized, one of them being less toxic to human cells than the naturally-derived starting materials. Structured-based inverted virtual screening and molecular dynamics simulations demonstrate that these active molecules likely target the insect odorant binding proteins and/or acetylcholinesterase and are able to form stable complexes. For the most promising compounds, nanoencapsulation assays were carried out in liposomes of egg phosphatidylcholine/cholesterol (7:3) prepared by both thin film hydration and ethanolic injection methods. The compound-loaded liposomes were generally monodisperse and with sizes smaller or around 200 nm. The thin film hydration method allowed high encapsulation efficiencies (above 85%) for both compounds and a delayed release, while for the systems prepared by ethanolic injection the encapsulation efficiency is lower than 50%, but the release is almost complete in two days.

Introduction

Crop destruction by pests, mainly by insects, is one of the main problems responsible for losses in agricultural production.[1-3] The strategy that has shown the best results in controlling insect pests is the use of synthetic pesticides, but they are associated with adverse impacts on the natural environment. Synthetic insecticides possess a broad spectrum of activity against various groups of insects, leading to almost complete elimination of the pests from the crops; nonetheless, the action of these compounds is not just limited to crop areas. Substantial amounts of insecticides applied in fields exert a toll in aquatic and land ecosystems, negatively impacting invertebrates and also vertebrates. Synthetic insecticides display significant environmental persistence, due to their large half-life, and have also propensity to accumulate in diverse trophic levels of the food net.[4-6]

The urgent need to adopt Integrated Pest Management to protect arable crop production, with the goal of decreasing the current overuse of synthetic insecticides, gives to plant products an important role.[7,8] For sustainable agriculture and public health, the future of pest management programs stresses on phytochemicals, either extracts or isolated molecules, as well as other biopesticides.[9,10] Bioinsecticides are associated with a number of advantages, namely lower toxicity to non-target organisms than synthetic, efficiency at low concentrations and ready biodegradability, which circumvents environmental pollution problems.[7,8,11,12]

Monoterpenes, secondary metabolites of plants are one of the most abundant and potent groups possessing biological activity against various pests.[13,14] Carvacrol (5-isopropyl-2-methylphenol)

[a.] *Centre of Chemistry, Department of Chemistry, University of Minho, Campus of Gualtar, 4710-057 Braga, Portugal. E-mail:msameiro@quimica.uminho.pt*
[b.] *REQUIMTE/LAQV, Laboratory of Pharmacognosy, Department of Chemistry, Faculty of Pharmacy, University of Porto, R. Jorge Viterbo Ferreira, 228, 4050-313 Porto, Portugal*
[c.] *Associate Laboratory i4HB - Institute for Health and Bioeconomy, Faculty of Medicine, University of Porto, 4200-319 Porto, Portugal;*
[d.] *UCIBIO – Applied Molecular Biosciences Unit, BioSIM - Department of Biomedicine, Faculty of Medicine, University of Porto, 4200-319 Porto, Portugal*
[e.] *Centre of Physics of Minho and Porto Universities (CF-UM-UP), University of Minho, Campus of Gualtar, 4710-057 Braga, Portugal*
Electronic Supplementary Information (ESI) available: ^{1}H and ^{13}C NMR spetra of all compounds and computational data. See DOI: 10.1039/x0xx00000x
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and thymol (2-isopropyl-5-methylphenol) are important phenolic monoterpenes obtained from the essential oil of some *Laminaceae* members, such as oregano, thyme, and savory, functioning as a chemical defence mechanism against phytopathogenic microorganisms.[15-21]

In addition to other monoterpenes, carvacrol and thymol have been structurally modified to enhance their biological activities.[22,23] For example, carvacrol, thymol and their derivatives have demonstrated larvicidal activity against *Aedes aegypti.*[RefA]Viviane Barros Silva The insecticidal action of *T. vulgaris* oil against *P.* shantungensis could be due to the presence of carvacrol and thymol as reported by Park *et al.* The authors also showed that thymol analogues have promising potential as first-choice insecticides against *P. shantungensis* adults and nymphs.[RefB] Jun-Hwan Park1, Ye Chitosan nanoparticles functionalized with β-cyclodextrin containing carvacrol and linalool revealed insecticidal activity against the species *Helicoverpa armigera* (corn earworm) and *Tetranychus urticae* (spider mite), as well as repellent activity and reduction in oviposition for the mites.[refC]

However, more studies are necessary to obtain other derivatives/analogues that can help to understand the structure-activity relationship, thus helping to plan structural changes and design of novel insecticides.

Considering all these facts and meeting the current interest in alternative botanical-based pesticides,[24,25] the present work is focused on the synthesis of a new set of carvacrol and thymol derivatives, including ether derivatives possessing the propyl chain without and with hydroxyl group and a chlorine atom as terminals, as well as the corresponding bicyclic ethers with an unsaturated chain containing a carboxylic acid terminal. The analogue bearing methoxyl, 1-hydroxyethyl and (3-chlorobenzoyl)oxy instead of the three original methyl groups present in carvacrol or thymol skeletons was also synthesized.

The insecticide activity of all compounds against the insect cell line *Sf9* (*Spodoptera frugiperda*) was evaluated. Moreover, computational studies were carried out to identity the most likely protein targets responsible for the observed insecticide activity of the new carvacrol and thymol derivatives. The two most active compounds were encapsulated in liposomes, considering the advantage in protecting them from early degradation and to surpass the high volatility generally observed in monoterpenes. Encapsulation efficiencies and release assays were carried out. Pursuing full biocompatibility, a natural phospholipid extract, phosphatidylcholine from egg yolk (Egg-PC) was chosen as main component for the formulation of liposomes, considering that phosphatidylcholines are major constituents of biological membranes. Liposomes composed of phosphatidylcholine (PC) and cholesterol (Ch) mixture, in a 7:3 ratio, are widely used as biological membrane models in studies of membrane permeation by antimicrobial peptides[26,27] and antitumor drug release.[28,29] Moreover, cholesterol is also known for its essential role as modulator of membrane fluidity. A promising utility for the encapsulated novel compounds is anticipated in future development of pesticide nanoformulations.

Results and discussion

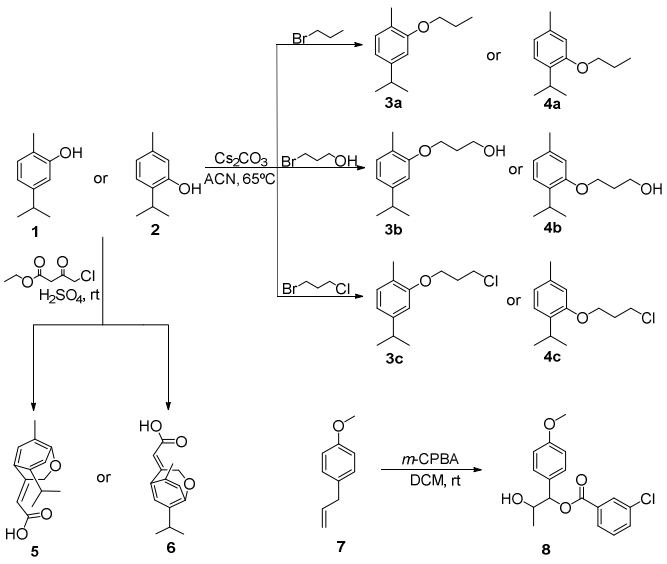
Synthesis

Essential oils components, such as carvacrol and thymol, can be sensitive to environmental conditions, such as temperature, light and oxygen, resulting in phenoxide radicals, which further proceed to oxidative, and polymerization reactions, that may lead to loss of insecticidal efficacy. They are safe and good starting materials allowing structural changes, namely through the hydroxyl group, and the obtention of derivatives can contribute to circumvent their stability limitations, improve physicochemical properties and result in compounds with improved biological activity. In this sense, carvacrol, 5-isopropyl-2-methylphenol **1** was reacted with 1-bromopropane, 3-bromopropan-1-ol and 1-bromo-3-chloropropane, using cesium carbonate as a base, by heating at 65 ºC in acetonitrile, followed by column chromatography on silica gel using dichloromethane/light petroleum as the eluent, to give the corresponding *O*-alkylated derivatives, namely 4-isopropyl-1-methyl-2-propoxybenzene **3a**, 3-(5-isopropyl-2-methylphenoxy)propan-1-ol **3b**, and 2-(3-chloropropoxy)-4-isopropyl-1-methylbenzene **3c** (Scheme 1). Although chemical structures of carvacrol and thymol isomers are very similar, only differing in the insertion of the OH-group placed in *ortho* (carvacrol) or *meta* (thymol) positions their activity may be different. Thus, starting from thymol, 2-isopropyl-5-methylphenol **2** and using again 1-bromopropane, 3-bromopropan-1-ol and 1-bromo-3-chloropropane in similar reaction conditions as mentioned above, 1-isopropyl-4-methyl-2-propoxybenzene **4a**, 3-(2-isopropyl-5-methylphenoxy)propan-1-ol **4b** and 2-(3-chloropropoxy)-1-isopropyl-4-methylbenzene **4c** were also obtained.

Compounds **3a**-**c** and **4a**-**c** were obtained as oils in 8 to 73% yields and were fully characterized by ^{1}H and ^{13}C NMR spectroscopy and HRMS. The ^{1}H NMR spectra of all compounds showed the methylenic protons of the ether-linked aliphatic chains introduced by alkylation reaction, namely $OCH_2CH_2CH_2$ as triplet (δ 3.99–4.38 ppm), $OCH_2CH_2CH_2$ as multiplet, triplet (**3b**) or quintet (**4c**) (δ 1.83-4.38 ppm) and $OCH_2CH_2CH_2$ as duplet (**3b**) or triplet (δ 3.80-4.09 ppm), as well as the methyl groups (δ 1.12-1.14 ppm, in **3a** and **4a**), in addition to the aromatic protons (δ 6.67-7.18 ppm) related to the carvacrol and thymol structures. ^{13}C NMR spectra showed the signals of methylenic carbons of the ether-linked aliphatic chains, namely $OCH_2CH_2CH_2$ (δ 64.11–69.36 ppm), $OCH_2CH_2CH_2$ (δ 22.65–32.56 ppm) and $OCH_2CH_2CH_2$ (δ 1.03–4.30 ppm), as well as the methyl carbons (δ ~10.60 ppm, in **3a** and **4a**), in addition to the remaining aromatic carbons (δ 109.32–157.15 ppm) related to carvacrol and thymol structures.

In order to obtain bicyclic ether derivatives possessing a conjugated system with simultaneously an aliphatic chain with a double linkage

and an electron withdrawing group, carvacrol **1** and thymol **2** were reacted with ethylchloroacetoacetate in acid medium at room temperature,[refD] followed by column chromatography purification in silica gel using dichloromethane/light petroleum as the eluent. Compounds **5** and **6** were obtained as solid materials in moderate yields and their structures were confirmed by the usual analytical techniques. ^{1}H NMR spectra showed the methylenic protons as doublets (δ 5.11 and 5.34 ppm), the double bond protons as triplets (δ 6.1 and 6.21 ppm), in addition to the remaining aliphatic protons for the dimethyl group (δ 1.24 and 1.15 ppm) and the isopropyl group protons (δ 3.10 – 3.22 ppm); the aromatic protons show up as singlets (δ 6.86 – 7.19 ppm). ^{13}C NMR spectra main features confirmed the presence of methylenic carbons (δ 73.45 and 72.42 ppm), and the double bond carbons (δ 116.08 and 112.32 ppm), in addition to the carbonyl group carbon (δ 174.39 and 174.29 ppm). Considering that a carboxylic ester as well as a chlorine atom in lateral chains of natural derivatives may increase their insecticidal activity, it was decided to prepare carvacrol/thymol analogues possessing (3-chlorobenzoyl)oxy, in addition to 1-hydroxyethyl and methoxyl groups instead of the three methyl original groups present in the considered isomers. Thus, reaction of anethole, 1-allyl-4-methoxybenzene **7** with *m*-chloroperbenzoic acid in dichloromethane gave (2-hydroxy-1-(4-methoxyphenyl)propyl 3-chlorobenzoate **8** as an oil in 89%. In the ^{1}H NMR spectrum stands out the protons of the 1-hydroxyethyl group as a multiplet (δ 4.19–4.26 ppm) and methoxyl group as a singlet (δ 3.81 ppm), as well as the aromatic protons of both rings as a series of six signals (δ 6.91 and 8.05 ppm). ^{13}C NMR spectrum showed the carbons of the 1-hydroxyethyl group (δ 70.20 ppm), the methoxyl group (δ 55.25 ppm), and carbonyl group (δ 164.71 ppm), as well as the aromatic carbons of both rings as 10 signals (δ 114.06 and 159.74 ppm).



Scheme 1. Synthesis of carvacrol and thymol derivatives/analogues **3a-c**, **4a-c**, **5**, **6** and **8**.

Screening of toxicity towards insect cells

Aiming the evaluation of the insecticidal activity of the synthesized carvacrol and thymol derivatives/analogues, *Spodoptera frugiperda* cells, a common pest, were used. For comparison purposes of their potency, all the molecules under study were screened at the same concentration (100 μg/mL). As can be seen in Figure 1, the naturally-derived starting materials carvacrol **1** and thymol **2**, displayed a marked toxicity towards *Sf9* insect cells, decreasing cell viability to *ca.* 55% and 40%, respectively. In a general way, the alkylation of the phenolic hydroxyl group of both carvacrol **1** and thymol **2** resulted in derivatives with lower toxicity (compounds **3a-c** and **4a-c**), pointing out the importance of this group to the activity displayed (Figure 1). Particularly, the *O*-alkylated derivatives containing propane and propan-1-ol (**3a,b** and **4a,b**) were completely devoid of toxicity. Noteworthy, when compared with **3a,b** and **4a,b**, the presence of a chlorine atom in the lateral chains of the natural derivatives significantly increased their toxicity (**3c** and **4c**). On the other hand, the syntheses of the bicyclic ether derivatives (compounds **5** and **6**) seem to lead to a reduction in toxicity, with a complete loss of activity in the case of compound **5**. However, compound **6** is still significantly active, with 50% of viability loss of *Sf9* cells (Figure 1). Among all derivatives/analogues synthesized, compounds **6** and **8** were the most potent, both eliciting *ca.* 50% of decrease in viability of insect cells (Figure 1).

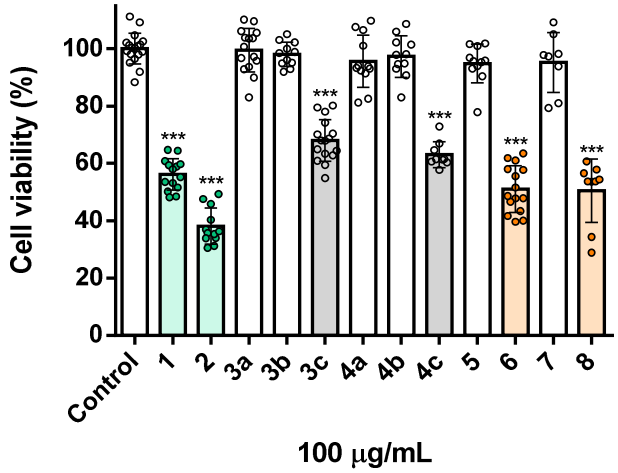


Figure 1. Viability of *Sf9* insect cells exposed to the molecules under study **3a-c**, **4a-c**, **5**, **6**, **7** and **8** (100 μg/mL), or medium (control). Cells were incubated for 24 h, after which viability was evaluated. *** $p < 0.001$.

Screening of toxicity towards human lung fibroblasts

Apart from their toxicity towards insect cells, we were also interested in evaluating the toxicity of these molecules in human cells, thus assessing their potential selectivity. Depending on how insecticides are used, they may affect human and environmental health in a variety of settings. Exposure to insecticides may occur through multiple routes, such as dermal contact, ingestion, and

inhalation, the latter being associated to the occupational or residential exposure during or after pesticides application, particularly on the form of aerosols. Thereby, as a model we used human lung fibroblasts (MRC-5 cell line), as they represent a cell population mimicking the major organ (lung) involved in pesticide exposure, namely *via* inhalation. As shown in Figure 2, the naturally derived starting materials carvacrol **1** and thymol **2**, displayed marked and equivalent toxicity towards MRC-5 cells, both decreasing cell viability to *ca.* 60%. This result indicates that the position of the hydroxyl group in the two molecules is irrelevant to the toxicity towards fibroblasts, contrary to what was verified against insect cells. On the other hand, the *O*-alkylated derivatives (**3a**-**c** and **4a**-**c**) were completely or nearly devoid of toxicity (Figure 2), derivatives holding a chlorine atom in the lateral chain (**3c** and **4c**) displaying a selective effect towards insect cells. It is also important to mention that compounds **6** and **8** were of equal or lower toxicity to MRC-5, when compared to the natural counterparts (Figure 2), respectively. Notably, compound **8** had a marginal effect in human cells (*ca.* 20% viability loss) (Figure 2) at the same concentration in which it caused *ca.* 50% of cell viability loss in insect cells (Figure 1), pointing out to its potential use as a semisynthetic insecticide.

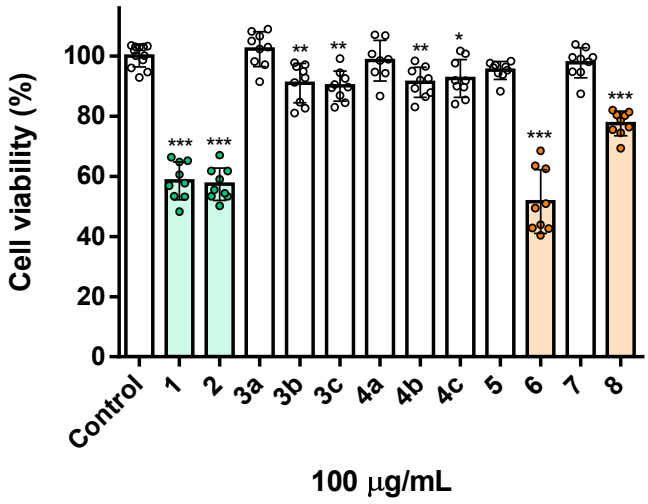


Figure 2: Viability of MRC-5 cells exposed to the molecules under study **3a-c**, **4a-c**, **5**, **6**, **7** and **8** (100 μg/mL), or medium (control). Cells were incubated for 24 h, after which viability was evaluated. * $p < 0.05$; ** $p < 0.01$; *** $p < 0.001$.

Nanoencapsulation and release assays

The most active compounds against *Sf9* cells, compounds **6** and **8**, were encapsulated in Egg-PC:Ch (7:3) liposomes, prepared by two methods, thin film hydration (TFH)[30] and ethanolic injection (EI),[31] and hydrodynamic sizes and polydispersity were measured by Dynamic Light Scattering (Table 1). Encapsulation efficiencies of both compounds were determined and are also presented in Table 1.

Table 1. Hydrodynamic diameter, polydispersity values and encapsulation efficiencies (EE%) (value ± standard deviation, SD, of three independent assays) of liposomal nanosystems of Egg-PC:Ch (7:3) containing compounds **6** and **8**.

Cpd	Method	Size ± SD (nm)	PDI ± SD	EE(%) ± SD(%)
6	EI	101.3 ± 22	0.24 ± 0.02	48.9 ± 0.5
	TFH	206.9 ± 28	0.26 ± 0.04	85.4 ± 0.8
8	EI	86.4 ± 19	0.21 ± 0.02	41 ± 6
	TFH	165.4 ± 15	0.15 ± 0.02	94.4 ± 2

It can be observed that the hydrodynamic sizes of compound-loaded nanosystems are around or below 200 nm and with a low polydispersity. Considering encapsulation efficiencies, the thin film hydration is the most suitable method for both compounds, exhibiting encapsulation efficiencies higher than 85% (attaining more than 94% for compound **8**). In the case of EI preparation method, the encapsulation efficiencies are lower than 50%.

The release of the encapsulated compounds **6** and **8** was followed for 48 h at room temperature towards a buffer of neutral pH (Figure 3).

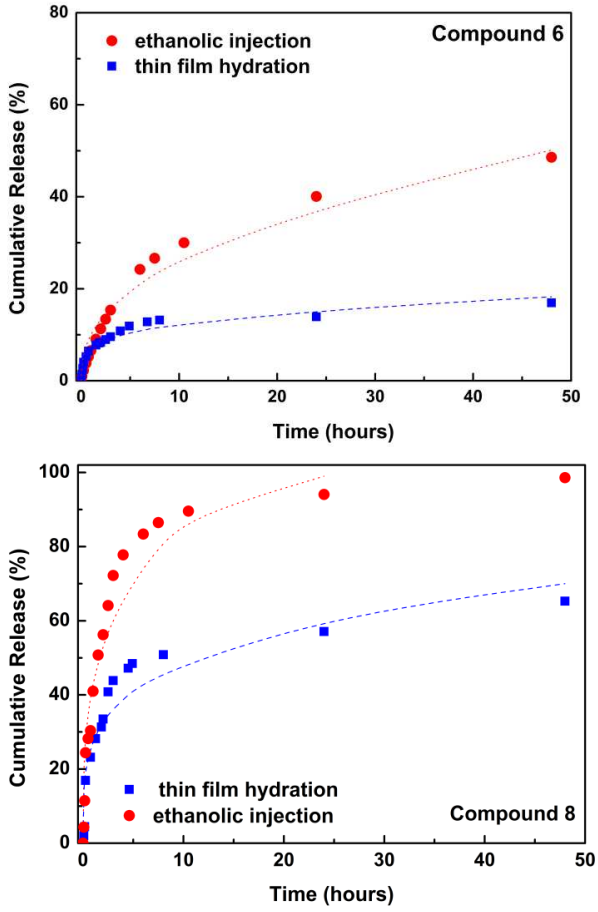


Figure 3. Release profiles and fitting to Korsmeyer-Peppas model. The lines are the fittings to the model.

The release profiles were fitted to the Korsmeyer-Peppas model[32] (fitted curves also in Figure 3), the results being presented in Table 2.

Table 2. Release parameters obtained by fitting to the Korsmeyer–Peppas model of the release profiles of compounds **6** and **8** from Egg-PC:Ch liposomal nanosystems. R is the coefficient of determination.

Cpd	Method	K (min^{-1})	n	R^2
6	EI	1.82×10^{-2}	0.416	0.98
	TFH	2.49×10^{-2}	0.250	0.97
8	EI	1.23×10^{-1}	0.305	0.97
	TFH	1.01×10^{-1}	0.252	0.95

The release is more effective from liposomes prepared with EI method, but the amount of compounds encapsulated is also significantly lower (considering EE% values). The liposomal systems obtained by TFH method exhibit a delayed release, with a much lower amount of compounds released in 48 h. In the case of liposomes loaded with compound **8** and prepared by EI method, the release was fitted only until 24 h (almost complete release at that time). Release of compound **8** is much faster than that of compound **6**, which seems to be retained in the liposomes. In all cases, the release mechanism is diffusion-controlled (n < 0.45).

Inverted Virtual Screening Results

Table S1 summarizes the average scores obtained for each protein-target with compounds **6** and **8**. GOLD scores are dimensionless, and a higher score means higher affinity, whereas Vina is the opposite. It uses a metric that is a more precise approximation of binding free energy, so a more negative value means better affinity. For each group of targets, the molecular structure with the highest score was selected and ranked from highest to lowest affinity according to the prediction of each scoring function. The scoring values obtained with the 5 independents scoring functions employed demonstrate that the most likely protein targets for the compounds **6** and **8** are Odorant Binding proteins (OBP), acetylcholinesterase (AChE), and chitinase. This tendency is consistent across the different methods evaluated.

Molecular Dynamics Simulations and Free Energy Calculations Results

Molecular dynamic simulations were performed considering the most likely proteins targets predicted in the inverted virtual screening protocol: AChE and OBPs. This was done to validate the results and evaluate the interactions between protein and ligand. The structures chosen were the ones presented the overall higher score: 1QON for AChE and 3K1E for OBPs.

The binding of compounds **6** and **8** towards these targets was evaluated, starting from the target-ligand complexes obtained from the docking calculations. The overall stability and strength of interaction was calculated through root mean square deviation (RMSD) for both the Cα atom of protein and ligand. The solvent accessible surface area (SASA) of the molecules was analysed, as well as the percentage of potential SASA of the ligands that was buried by the target upon binding, and the number of hydrogen bonds formed throughout time. A summary of all these results is presented in Table 3.

Analysis of the RMSD values presented in Figure S1 shows that compounds **6** and **8** retain the overall positions predicted from docking along the 100 ns of MD simulation performed (average RMSD values between 0.6 and 1.5 Å). In addition, an analysis of the percentage of the potential SASA of the ligand buried by the target (Table 3 and Figure S2) demonstrates that the protein-ligand complexes remain quite stable throughout the simulations, remaining well-shielded from the solvent by AChE and OBP. Compound **8** remains strongly associated to OBP, maintaining on average 96% of its surface non-solvent accessible. With AChE, compound **8** has only 87% of its surface protected by the protein. Compound **6** remains also strongly associated to OBP and AChE with average values of non-solvent accessible surface of 87 and 89%, respectively.

Analysing the number of hydrogen bonds formed between each molecule and their putative target throughout the simulation is helpful to understand the strength and type of interactions formed (Table 3 and Figure S3). Compound **6** establishes on average more hydrogen bonds with both targets than compound **8**. This effect is more evident for AchE (2 hydrogen bonds on average for compound **6** *vs* 0.5 for compound **8**).

Table 3 also presents the values for the overall Gibbs energy of association calculated using MM-GBSA. The results show that compound **8** has stronger affinity towards both targets than compound **6**, with binding free energies ranging from -32.5 and -32.2 kcal/mol for AChE and OBP, against -21.3/-19.4 kcal/mol. As exhibited by the MM-GBSA values, compounds **6** and **8** display very similar preferences for AChE and OBP.

OBP are a class of proteins that are also present in many different organisms, from mammals to invertebrates. Their function is to bind and transport small hydrophobic molecules into the olfactory receptors. In insects they are present in high number, but they present common features such as their small size, six alpha-helix domains and six cysteine residues bound by three disulfide bonds.[33-35]

Table 3. Average protein and ligand RMSD values (Å), Average complex SASA (Å) and percentage of ligand SASA Buried (%), number of average hydrogen bonds formed between the targets and ligands obtained for AChE and OBP complexes. ΔG binding energy was determined using MM/GBSA with depiction of the most important residues.

	Compound	Average RMSD of the ligand (Å)	SASA (Å^2)	Percentage of ligand SASA buried (%)	Average H-bonds	ΔGbind (kcal/mol)	Main contributors
AChE	**6**	0.7 ± 0.2	45.5 ± 14.8	89	1.5 ± 0.8	-19.4 ± 0.2	Arg70 (-6.4 ± 3.2) Tyr374 (-2.5 ± 2.0) Trp83 (-2.2 ± 0.5)
	8	1.1 ± 0.2	68.7 ± 18.2	87	0.5 ± 0.6	-32.2 ± 0.2	Trp83 (-2.8 ± 0.5) Tyr374 (-1.9 ± 0.7) Tyr71 (-1.7 ± 0.9)
OBP	**6**	0.6 ± 0.3	52.1 ± 19.2	87	0.7 ± 0.9	-21.3 ± 0.5	Arg85 (-3.1 ± 4.4) Ser108 (-2.8 ± 2.0) Trp105 (-2.5 ± 0.7)
	8	1.5 ± 0.4	19.2 ± 8.6	96	0.3 ± 0.5	-32.5 ± 0.2	Phe114 (-1.6 ± 1.1) Leu71 (-1.5 ± 0.5) Trp105 (-1.2 ± 0.5)

There is a virtual screening study regarding the effect of several essential oils as pesticide agents, with promising results with thymol, carvacrol and OBP, when compared with DEET, a compound commonly used as mosquito repellent.[36] There is also deposited in the Protein Data Bank, a structure of a mammal OBP bound to thymol (PDB: 1E02). While the structure is different from insect OBP, it maintains a similar structural function of binding to small odorant molecules. Both compounds **6** and **8** are lipophilic molecules but compound **8** seems to have a higher binding affinity toward OBP.

When bound to OBP1, compound **6** is stabilized primarily by electrostatic interactions with Arg85 (-3.1 ± 4.4), Ser108 (-2.8 ± 2.0) and Trp105 (-2.5 ± 0.7). Compound **8** is stabilized mainly by non-polar interactions with Trp105 (-1.2 ± 0.5), Phe114 (-1.6 ± 1.1) and Leu71 (-1.5 ± 0.5) (Figure 4). The results seem to suggest that compound **8** can be a good antagonist for OBP.

AChE is a common target for several pesticides, due to its crucial role in the hydrolysis of acetylcholine. This serine hydrolase is a common neurotransmitter regulator in many species, from mammals to insects, hence, the lack of specificity of pesticides directed to this target lead to a series of health and environmental problems.[37,38] Structural studies with AChE led to the discovery that the pesticides worked by phosphorylating a conserved serine residue in the active site. However, targeting specific structural differences between AChE of different species can be a powerful tool in the discovery of new and more specific pesticides.[37]

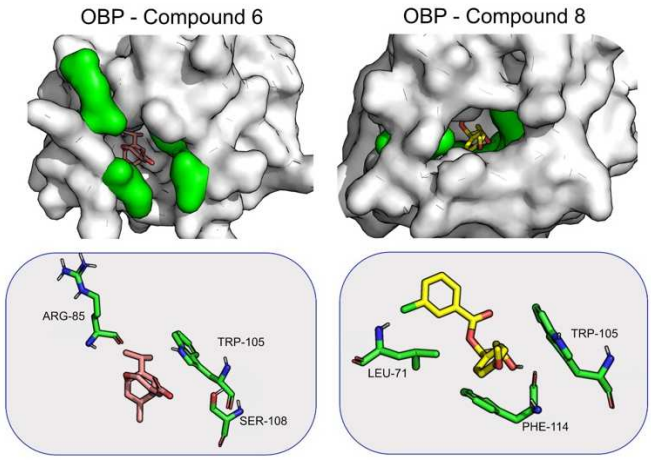


Figure 4. OBP - compound 6 (pink licorice) and compound 8 (yellow licorice) interaction maps. Blue arrows represent π-π stacking and red lines represent hydrogen bonds. The three most relevant residues are represented in green licorice.

To assess the specificity of these two molecules towards insect AChE in comparison with human AChE, docking studies were performed to the human form of AChE (PDB: 5HFA) with all the scoring functions tested in this study (data shown in Supplementary Information, Table S2). Relevantly, the results consistently showed that compounds **6** and **8** presented higher docking scores against insect AChE than towards human AChE. This tendency was observed with the five independent scoring functions evaluated, suggesting that these molecules show strong preference to insect AChE binding in comparison to human AChE.

When bound to insect AChE (1QON), compound **6** is stabilized mainly by Arg70 (-6.4 ± 3.2), Tyr374 (-2.5 ± 2.0) and Trp83 (-2.2 ± 0.5) by hydrogen bonds and electrostatic interactions. For compound **8**, the stabilization mainly comes from π-π interactions, as well as hydrogen bond formed with Tyr374. Hydrogen bonds are formed also with Trp83 (-1.9 ± 0.4) (Figure 5).

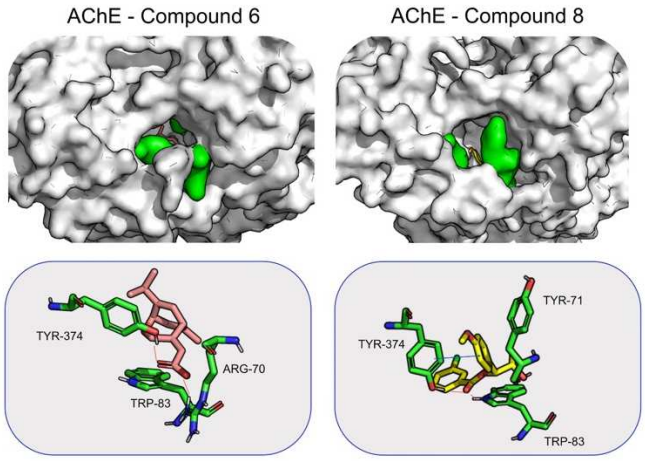


Figure 5. AChE - compound 6 (pink licorice) and compound 8 (yellow licorice) interaction maps. Blue arrows represent π-π stacking and red lines represent hydrogen bonds. The three most relevant residues are represented in green licorice.

Experimental

Chemistry

TLC analyses were carried out on 0.25 mm thick precoated silica plates (Merck Fertigplatten Kieselgel 60F254) and spots were visualized under UV light. Chromatography on silica gel was carried out on Merck Kieselgel (230-240 mesh). NMR spectra were obtained on a Bruker Avance III at an operating frequency of 400 MHz for ^{1}H NMR and 100.6 MHz for ^{13}C NMR using the solvent peak as internal reference at 25 ºC. All chemical shifts are given in ppm using δ Me_4Si = 0 ppm as reference and *J* values are given in hertz. Assignments were made by comparison of chemical shifts, peak multiplicities and *J* values and were supported by spin decoupling-double resonance and bidimensional heteronuclear correlation techniques. High resolution mass spectrometry analyses were performed at the "CACTI – Centro de Apoio Científico-Tecnolóxico á Investigación", at University of Vigo, Spain.

General procedure for the synthesis of compounds 3a-c and 4a-c

To a solution of carvacrol, 5-isopropyl-2-methylphenol **1** (1 equiv), thymol, 2-isopropyl-5-methylphenol **2** (1 equiv) in acetonitrile (4 mL), the corresponding alkyl halide (1.1 equiv) and cesium carbonate (5 equiv) were added, and the resulting mixture was heated at 65 ºC for 2 h 30 min. The progress of the reaction was monitored by TLC (light petroleum). The excess of base was filtered, the solvent was evaporated and the crude mixture was purified by column chromatography on silica gel using dichlorometane/light petroleum (mixtures of increasing polarity) as the eluent.

4-Isopropyl-1-methyl-2-propoxybenzene 3a. Starting from carvacrol, 5-isopropyl-2-methylphenol **1** (0.105 mL, 6.7×10^{-3} mol) and using 1-bromopropane (0.067 mL, 7.3×10^{-3} mol), compound **3a** was obtained as a light orange oil (0.069g, 54% yield). R*f* = 0.56 (light petroleum). ^{1}H NMR ($CDCl_3$, 400 MHz): δ_H 1.12 (3H, t, *J* 6.8 Hz, $OCH_2CH_2CH_3$), 1.30 (6H, d, *J* 6.8 Hz, $CH(CH_3)_2$), 1.84 – 1.93 (2H, m, $OCH_2CH_2CH_3$), 2.26 (3H, s, CH_3Ph), 2.88-2.95 (1H, m, $CH(CH_3)_2$), 3.99 (2H, t, *J* 6.4 Hz, $OCH_2CH_2CH_3$), 6.75 (1H, d, *J* 1.2 Hz, H-3), 6.78 (1H, dd, *J* 7.6 and 1.6 Hz, H-5), 7.11 (1H, d, *J* 7.2 Hz, H-6) ppm. ^{13}C NMR ($CDCl_3$, 100.6 MHz): δ_C 10.68 ($OCH_2CH_2CH_3$), 15.77 (CH_3Ph), 22.80 ($OCH_2CH_2CH_3$), 24.13 ($CH(CH_3)_2$), 34.14 ($CH(CH_3)_2$), 69.36 ($OCH_2CH_2CH_3$), 109.45 (C-3), 117.76 (C-5), 124.13 (C-1), 130.32 (C-6), 147.79 (C-4), 157.15 (C-2) ppm. HRMS: *m/z* (ESI-TOF): Calcd. for $C_{13}H_{21}O$ $[M+1]^+$ 193.1587; found 193.1592.

3-(5-Isopropyl-2-methylphenoxy)propan-1-ol 3b. Starting from carvacrol, 5-isopropyl-2-methylphenol **1** (0.105 mL, 6.6×10^{-3} mol) and using 3-bromopropan-1-ol (0.065 mL, 7.3×10^{-3} mol), compound **3b** was obtained as a colorless oil (0.011g, 8% yield). R*f* = 0.40 (dichloromethane/light petroleum 1:1). ^{1}H NMR ($CDCl_3$, 400 MHz): δ_H 1.24 (6H, d, *J* 6.8 Hz $CH(CH_3)_2$), 2.15 – 2.22 (5H, m, CH_3Ph and $OCH_2CH_2CH_2OH$), 2.81 – 2.91 (1H, m, $CH(CH_3)_2$), 4.09 (2H, t, *J* 6.0 Hz, $OCH_2CH_2CH_2OH$), 4.38 (2H, t, *J* 6.0 Hz, $OCH_2CH_2CH_2OH$), 6.69 (1H, d, *J* 1.6 Hz, H-6), 6.74 (1H, dd, *J* 7.6 and 1.6 Hz, H-4), 7.06 (1H, d, *J* 7.6 Hz, H-3) ppm. ^{13}C NMR ($CDCl_3$, 100.6 MHz): δ_C 15.76 (CH_3Ph), 24.11 ($CH(CH_3)_2$), 28.85 ($OCH_2CH_2CH_2Cl$), 34.12 ($CH(CH_3)_2$), 63.77 ($OCH_2CH_2CH_2OH$), 64.93 ($OCH_2CH_2CH_2OH$), 109.32 (C-6), 118.18 (C-4), 124.08 (C-2), 130.43 (C-3), 147.89 (C-5), 156.68 (C-1) ppm. HRMS: *m/z* (ESI-TOF): Calcd. for $C_{13}H_{21}O_2$ $[M+1]^+$ 209.1536; found 209.1539.

2-(3-Chloropropoxy)-4-isopropyl-1-methylbenzene 3c. Starting from carvacrol, 5-isopropyl-2-methylphenol **1** (0.210 mL, 1.3×10^{-3} mol) and using 1-bromo-3-chloropropane (0.140 mL, 1.4×10^{-3} mol), compound **3c** was obtained as a colorless oil (0.175 g, 57% yield). R*f* = 0.29 (light petroleum). ^{1}H NMR ($CDCl_3$, 400 MHz): δ_H 1.26 (6H, d, *J* 6.8 Hz, $CH(CH_3)_2$), 2.20 (3H, s, CH_3Ph), 2.25 – 2.31 (2H, m, $OCH_2CH_2CH_2Cl$), 2.83 – 2.94 (1H, m, $CH(CH_3)_2$), 3.80 (2H, t, *J* 6.4 Hz, $OCH_2CH_2CH_2Cl$), 4.14 (2H, t, *J* 6.0 Hz, $OCH_2CH_2CH_2Cl$), 6.73 (1H, d, *J* 1.6 Hz, H-3), 6.77 (1H, dd, *J* 7.6 and 1.6 Hz, H-5), 7.08 (1H, d, *J* 7.6 Hz, H-6) ppm. ^{13}C NMR ($CDCl_3$, 100.6 MHz): δ_C 15.76 (CH_3Ph), 24.11 ($CH(CH_3)_2$), 32.56 ($OCH_2CH_2CH_2Cl$), 34.13 ($CH(CH_3)_2$), 41.70 ($OCH_2CH_2CH_2Cl$), 64.16 ($OCH_2CH_2CH_2Cl$), 109.49 (C-3), 118.27 (C-5), 124.07 (C-1), 130.46 (C-6), 147.96 (C-4), 156.66 (C-2) ppm. HRMS: *m/z* (ESI-TOF): Calcd. for $C_{13}H_{20}{}^{35}ClO$ $[M+1]^+$ 227.1197; found 227.1198; Calcd. for $C_{13}H_{20}{}^{37}ClO$ $[M+1]^+$ 229.1171; found 229.1170.

1-Isopropyl-4-methyl-2-propoxybenzene 4a. Starting from thymol, 2-isopropyl-5-methylphenol **2** (0.105 g, 7.0 × 10^{-3} mol) and using 1-bromopropane (0.122 mL, 7.3 × 10^{-3} mol), compound **4a** was obtained as a light yellow oil (0.098g, 73% yield). R*f* = 0.78 (light petroleum). ^{1}H NMR ($CDCl_3$, 400 MHz): δ_H 1.14 (3H, t, *J* 7.2 Hz, $OCH_2CH_2CH_3$), 1.29 (6H, d, *J* 7.2 Hz, $CH(CH_3)_2$), 1.83 – 1.96 (2H, m, $OCH_2CH_2CH_3$), 2.39 (3H, s, CH_3Ph), 3.33 – 3.44 (1H, m, $CH(CH_3)_2$), 3.99 (2H, t, *J* 6.4 Hz, $OCH_2CH_2CH_3$), 6.74 (1H, d, *J* 1.6 Hz, H-3), 6.80 (1H, dd, *J* 7.6 and 1.6 Hz, H-5), 7.16 (1H, d, *J* 7.6 Hz, H-6) ppm. ^{13}C NMR ($CDCl_3$, 100.6 MHz): δ_C 10.60 ($OCH_2CH_2CH_3$), 21.10 (CH_3Ph), 22.54 ($CH(CH_3)_2$), 22.65 ($OCH_2CH_2CH_3$), 26.48 ($CH(CH_3)_2$), 69.19 ($OCH_2CH_2CH_3$), 112.01 (C-3), 120.67 (C-5), 125.60 (C-6), 133.82 (C-1), 135.99 (C-4), 156.02 (C-2) ppm. HRMS: *m/z* (ESI-TOF): Calcd. for $C_{13}H_{21}O$ $[M+H]^+$ 193.1587; found 193.1586.

3-(2-Isopropyl-5-methylphenoxy)propan-1-ol 4b. Starting from thymol, 2-isopropyl-5-methylphenol **2** (0.106 g, 7.0 × 10^{-3} mol) and using 3-bromopropan-1-ol (0.065 mL, 7.3 × 10^{-3} mol), compound **4b** was obtained as a light yellow oil (0.019g, 13% yield). R*f* = 0.47 (dichloromethane/light petroleum 1:1). ^{1}H NMR ($CDCl_3$, 400 MHz): δ_H 1.21 (6H, d, *J* 7.2 Hz, $CH(CH_3)_2$), 2.16 – 2.23 (2H, m, $OCH_2CH_2CH_2OH$), 2.33 (3H, s, CH_3Ph), 3.23 – 3.33 (1H, m, $CH(CH_3)_2$), 4.07 (2H, t, *J* 6.0 Hz, $OCH_2CH_2CH_2OH$), 4.38 (2H, t, *J* 6.4 Hz, $OCH_2CH_2CH_2OH$), 6.67 (1H, d, *J* 1.6 Hz, H-6), 6.76 (1H, dd, *J* 8.0 and 1.6 Hz, H-4), 7.10 (1H, d, *J* 8.0 Hz, H-3) ppm. ^{13}C NMR ($CDCl_3$, 100.6 MHz): δ_C 21.30 (CH_3Ph), 22.73 ($CH(CH_3)_2$), 26.54 ($CH(CH_3)_2$), 28.87 ($OCH_2CH_2CH_2Cl$), 63.80 ($OCH_2CH_2CH_2OH$), 64.94 ($OCH_2CH_2CH_2OH$), 112.11 (C-6), 121.24 (C-4), 125.87 (C-3), 134.00 (C-2), 136.28 (C-5), 155.67 (C-1) ppm. HRMS: *m/z* (ESI-TOF): Calcd. for $C_{13}H_{21}O_2$ $[M+H]^+$ 209.1536; found 209.1530.

2-(3-Chloropropoxy)-1-isopropyl-4-methylbenzene 4c. Starting from thymol, 2-isopropyl-5-methylphenol **2** (0.254 g, 1.7 × 10^{-3} mol) and using 1-bromo-3-chloropropane (0.180 mL, 1.8 × 10^{-3} mol), compound **4c** was obtained as a colorless oil (0.116g, 30% yield). R*f* = 0.28 (light petroleum). ^{1}H NMR ($CDCl_3$, 400 MHz): δ_H 1.29 (6H, d, *J* 7.2 Hz, $CH(CH_3)_2$), 2.33 (2H, quint, *J* 6.4 Hz, $OCH_2CH_2CH_2Cl$), 2.40 (3H, s, CH_3Ph), 3.30 – 3.40 (1H, m, $CH(CH_3)_2$), 3.84 (2H, t, *J* 6.4 Hz, $OCH_2CH_2CH_2Cl$), 4.18 (2H, t, *J* 5.6 Hz, $OCH_2CH_2CH_2Cl$), 6.76 (1H, d, *J* 1.6 Hz, H-3), 6.83 (1H, dd, *J* 7.6 and 1.6 Hz, H-5), 7.18 (1H, d, *J* 7.6 Hz, H-6) ppm. ^{13}C NMR ($CDCl_3$, 100.6 MHz): δ_C 21.27 (CH_3Ph), 22.72 ($CH(CH_3)_2$), 26.57 ($CH(CH_3)_2$), 32.49 ($OCH_2CH_2CH_2Cl$), 41.67 ($OCH_2CH_2CH_2Cl$), 64.11 ($OCH_2CH_2CH_2Cl$), 112.19 (C-3), 121.31 (C-5), 125.86 (C-6), 133.93 (C-1), 136.29 (C-4), 155.63 (C-2) ppm. HRMS: *m/z* (ESI-TOF): Calcd. for $C_{13}H_{20}{}^{35}ClO$ $[M+H]^+$ 227.1197; found 227.1196; Calcd. for $C_{13}H_{20}{}^{37}ClO$ $[M+H]^+$ 229.1171; found 229.1164.

General procedure for the synthesis of compounds 5 and 6

To a solution of carvacrol, 5-isopropyl-2-methylphenol **1** (1 equiv) or thymol, 2-isopropyl-5-methylphenol **2** (1 equiv) in 70% aqueous sulfuric acid (5 mL), ethylchloroacetoacetate (1.5 equiv) was added and kept under stirring at room temperature for 5 or 3 days, respectively. The reaction mixture was poured into ice water and stirred for 2 h to give a precipitate. The solid was collected by filtration, washed with cold water, dried in a vacuum oven, and the crude mixture was purified by column chromatography on silica gel using dichloromethane/light petroleum 90:10, as the eluent.

7-Methyl-9-(propan-2-yl)-2-oxabicyclo[3.2.2]nona-1(7),5,8-trien-4-ylidene)ethanoic acid 5. Starting from carvacrol, 5-isopropyl-2-methylphenol **1** (0.100 g, 0.67 × 10^{-3} mol) and ethylcholoroacetoacetate (0.138 mL, 1.0 × 10^{-3} mol), compound **5** was obtained as a light yellow solid (0.048 g, 29% yield). R_f = 0.20 (dichloromethane). ^{1}H NMR ($CDCl_3$, 400 MHz): δ_H 1.24 (6H, d, *J* 6.8 Hz, $CH(CH_3)_2$), 2.25 (3H, s, $PhCH_3$), 3.11 – 3.22 (1H, m, $CH(CH_3)_2$), 5.11 (2H, d, *J* 2.0 Hz, CCH_2O), 6.10 (1H, t, *J* 1.6 Hz, $CHCO_2H$), 6.86 (1H, s, Ar-H), 6.97 (1H, s, Ar-H) ppm. ^{13}C NMR ($CDCl_3$, 100.6 MHz): δ_C 15.25 ($PhCH_3$), 24.07 ($CH(CH_3)_2$), 29.36 ($CH(CH_3)_2$), 73.45 (CCH_2O), 113.11 (Ar-C), 116.08 ($CHCO_2H$), 121.41 (Ar-Cq), 121.90 (Ar-Cq), 130.16 (Ar-C), 147.64 (Ar-Cq), 156.11 (Ar-Cq), 164.48 ($C(CH_2)$) 174.39 (CO_2H) ppm. HRMS: *m/z* (ESI-TOF): Calcd. for $C_{14}H_{17}O_3$ $[M+H]^+$ 233.1172; found 233.1180.

9-Methyl-7-(propan-2-yl)-2-oxabicyclo[3.2.2]nona-1(7),5,8-trien-4-ylidene)ethanoic acid 6. Starting from thymol, 2-isopropyl-5-methylphenol **2** (0.256 g, 1.01 × 10^{-3} mol) and ethylchloroacetoacetate (0.337 mL, 2.50 × 10^{-3} mol), compound **6** was obtained as a light yellow solid (0.051 g, 20% yield). R*f* = 0.91 (dichloromethane/methanol, 95:5). ^{1}H NMR (DMSO-d_6, 400 MHz): δ_H 1.15 (6H, d, *J* 6.8 Hz, $CH(CH_3)_2$), 2.32 (3H, s, $PhCH_3$), 3.10 – 3.18 (1H, m, $CH(CH_3)_2$), 5.34 (2H, d, *J* 1.6 Hz, CCH_2O), 6.21 (1H, t, *J* 1.6 Hz, $CHCO_2H$), 6.73 (1H, s, Ar-H), 7.19 (1H, s, Ar-H), 9.93 (1H, s, OH) ppm. ^{13}C NMR (DMSO-d_6, 100.6 MHz): δ_C 21.68 (CH_3), 22.26 ($CH(CH_3)_2$), 26.33 ($CH(CH_3)_2$), 72.42 (CCH_2O), 112.32 ($CHCO_2H$), 118.11 (Ar-C), 120.37 (Ar-C), 126.40 (Ar-C), 132.36 (Ar-C), 136.83 (Ar-C), 156.74 (Ar-C), 164.55 ($C(CH_2)$) 174.29 (CO_2H) ppm. HRMS: *m/z* (ESI-TOF): Calcd. for $C_{14}H_{17}O_3$ $[M+H]^+$ 233.1172; found 233.1177.

Synthesis of (2-hydroxy-1-(4-methoxyphenyl)propyl 3-chlorobenzoate 8.

To a suspension of *m*-CPBA (55%; 1.261 g, 7.31 mmol, 3.0 equiv) and sodium hydrogen carbonate (0.343 g, 4.08 mmol, 3.0 equiv) in dichloromethane (2.8 mL) at 0 °C (ice bath), with stirring, a solution of anethole, 1-allyl-4-methoxybenzene **7** (0.2 mL, 1.35 mmol) in dichloromethane (2 mL) was added. The reaction mixture was allowed to stir at room temperature for 23 h. Then, it was washed with aqueous solution of sodium sulfite 10% (2×5 mL) and a saturated aqueous solution of sodium hydrogen carbonate (2×5 mL). The organic phase was dried with magnesium sulfate and the solvent was evaporated on a rotary evaporator, yielding compound **8** as a yellow oil (0.384 g, 89%). R_f = 0.40 (dichloromethane/methanol, 99:1). ^{1}H NMR ($CDCl_3$, 400 MHz): δ_H

1.15 (3H, d, *J* 6.4 Hz, CH_3), 3.81 (3H, s, OCH_3), 4.19 – 4.26 (1H, m, C*H*(CH_3)OH), 5.75 (1H, d, *J* 7.6 Hz, C*H*CO_2PhCl), 6.78 (1H, s, OH), 6.91 (2H, d, *J* 8.8 Hz, H-3 and H-5 PhOCH_3), 7.35 (2H, d, *J* 8.8 Hz, H-2 and H-6 PhOCH_3), 7.40 (1H, t, *J* 8.0 Hz, H-5 PhCl), 7.55 – 7.57 (1H, m, H-6 PhCl), 7.94 – 7.98 (1H, m, H-4 PhCl), (8.05 (1H, d, *J* 1.6 Hz, H-2 PhCl) ppm. ^{13}C NMR $CDCl_3$, 100.6 MHz): δ_C 18.85 (CHCH*C*H_3), 55.25 (OCH_3), 70.20 (CH*C*HCH_3), 81.42 (*C*HCHCH_3), 114.06 (Ar-C-2 and Ar-C-6 PhOCH_3), 127.83 (Ar-C-3 or Ar-C-5 PhOCH_3), 128.52 (Ar-C-5 or Ar-C-3 PhOCH_3), 129.32 (Ar-C-4 PhOCH_3), 129.64 (Ar-C-6 PhCl), 129.75 (Ar-C-2 and Ar-C-5 PhCl), 131.86 (Ar-C-1 PhCl), 133.14 (Ar-C-4 PhCl), 134.56 (Ar-C-3 PhCl), 159.74 (Ar-C-1 PhOCH_3), 164.71 (C=O) ppm. HRMS: *m/z* (ESI-TOF): Calcd. for $C_{17}H_{17}{}^{35}ClO_4$ $[M+Na]^+$ 343.0711; found 343.0708; Calcd. for $C_{17}H_{17}{}^{37}ClO_4$ $[M+Na]^+$ 345.0684; found 345.0679.

Cell culture

Sf9 (*Spodoptera frugiperda*) cells were maintained as a suspension culture and cultivated in Grace's medium with 10% FBS and 1% penicillin/streptomycin, at 28 ºC. Cells were used in experiments while in the exponential phase of growth. On the other hand, MRC-5 (human lung fibroblasts) cells were culture in MEM supplemented with 10% FBS and 1% penicillin/streptomycin at 37 ºC, in a humidified atmosphere of 5% CO_2.

Viability assessment

For the assessment of viability, a resazurin-based method was used, similarly to what we described before.[24] *Sf9* and MRC-5 cells were plated at a density of 3.0×10^4 and 2.0×10^4 cells/well, respectively, incubated for 24 h and then exposed to the molecules under study for 24 h. After this period, a commercial solution of resazurin was added (1:10) and the kinetic reaction of fluorescence increase monitored at 560/590 nm. For both cell lines, 60 min. of incubation were used.

Nanoencapsulation studies

Egg yolk phosphatidylcholine (Egg-PC) and cholesterol (Ch) were obtained from Sigma-Aldrich. In the ethanolic injection method (EI),[31] nanoliposomes were prepared by injection of an ethanolic solution of lipids/compounds mixture (in the ratio Egg-PC:Ch 7:3) in an aqueous buffer solution under vigorous stirring. In the thin film hydration (TFH) method,[30] a thin film of the Egg-PC/Ch mixture was obtained evaporating a lipid solution in chloroform under an ultrapure nitrogen stream. The compound solution was added and, after evaporation, the film was hydrated with the aqueous solution, followed by sonication and ten extrusion cycles through polycarbonate membranes (5×400 nm and 5×200 nm) in a Lipex™ extruder (from Northern Lipids). In both methods, the final lipid concentration was 1 mM.

The encapsulation efficiency (percent), *EE%*, was determined through absorbance measurements. After preparation, liposomes were subjected to centrifugation in Amicon® Ultra centrifugal filter units 100 kDa at 11.000 rpm for 60 min. Then, the supernatant was removed and its absorption spectrum was measured in a Shimadzu UV-3600 Plus UV/Vis/NIR spectrophotometer. Using a previously measured calibration curve of absorbance *versus* concentration, the encapsulation efficiencies of both compounds were determined through equation (1), and three independent assays were performed.

$$EE(\%) = \frac{Total\ quantity - Quantity\ of\ non{-}encapsulated\ compound}{Total\ quantity} \times 100 \qquad (1)$$

The compound-loaded liposomes' mean hydrodynamic diameter and size distribution (polydispersity) were measured with a Dynamic Light Scattering (DLS) equipment Litesizer 500 from Anton Paar, at 25 °C, using a solid-state laser of 648 nm and 40 mW. Five independent measurements were carried out for each of the samples.

Release assays to phosphate buffer (pH=7) were performed during 48 h in triplicate, using dialysis membranes. Calibration curves of absorbance *versus* concentration, previously determined for each compound, were used to calculate the concentration of released compounds. The loaded liposomes were kept under stirring at 25 °C, the solutions being covered to prevent evaporation. The Korsmeyer-Peppas model[32] was used to describe the compound release kinetics from the liposomes (equation 2):

$$\frac{M_t}{M_\infty} = K \cdot t^n \qquad (2)$$

where $\frac{M_t}{M_\infty}$ represents the fraction of released drug, *K* is the release constant, *n* the transport exponent (dimensionless) and *t* is the time. The transport exponent is directly related to the release mechanism of the compound: if $n > 1$, the release is controlled by swelling and material relaxation; $0.89 < n < 1$ indicates a relaxation-controlled mechanism, $0.45 < n < 0.89$ indicates a combination of diffusion and erosion in drug release (non-Fickian release) and when $n < 0.45$ the release mechanism is diffusion-controlled (Fickian release).

Molecular Docking and Inverted Virtual Screening optimization

A selection of putative molecular targets associated to insecticide activity was made and was used as a basis for the design of an inverted virtual screening protocol to identity the most likely protein targets responsible for the observed insecticide activity of the carvacrol and thymol derivatives evaluated. The selection of putative molecular targets was made by scanning Scopus using the keywords "Virtual Screening" and "biopesticides". Some targets were selected from the Protein Data Bank (PDB) based upon the relevance of the target. The relevance and year of publication were considered for the final selection. Of the seventeen studies analyzed, 23 PDB structures were selected to proceed to the optimization of the Inverted Virtual Screening protocol. The molecular targets under consideration are detailed in Table S3.

The 23 molecular targets extracted from the PDB were prepared using the Pymol Autodock Vina plugin[56] where crystallographic waters and ligands (when present) were removed. The crystallographic ligands were then saved in separate files and used as reference for active site coordinates, as well as validation for the re-docking steps. In the absence of ligands, the active site coordinates were based on the most important residues described in the literature. Re-docking was used as a validation tool, measuring the ability of the docking software in reproducing the orientation of the crystallographic pose.

Five docking scoring functions were used, and the protocol was optimized for each one and customized to each molecular target with the goal of minimizing the RMSD between the crystallographic and docking poses at the re-docking stage. GOLD[57] (PLP, ASP, ChemScore, and GoldScore scoring function) and Autodock Vina[58] were the selected docking software as they are well established and are widely used in research.[59,60] While it may be difficult to compare different scoring functions, as they handle the target and ligand in different manners, hence different metrics and scales,[61] testing multiple alternatives eliminates bias and enables the evaluation of clear interaction tendencies.

Molecular Dynamics Simulations and Free Energy Calculations

Two of the compounds that presented higher insecticide activity when tested experimentally (compounds **6** and **8**) were simulated in complex with the two most promising targets identified from the inverted virtual screening study (acetylcholinesterase - 1QON and odorant binding protein 1 - 3KIE). Molecular dynamics simulations were performed using the Amber18 software.[62]

For MD simulations, the target structures cannot present any gaps. Therefore, a homology model was created for the structure 1QON using SWISS-MODEL[63] (detail in Supplementary Information, Figure S4).

The complexes selected for the MD simulations were the ones that were predicted during the inverted virtual screening experiments using GOLD/PLP and were subsequently treated with the Leap module of AMBER.[64] The protein targets were treated with the ff14SB force field,[65] while the compounds **6** and **8** were parameterized using ANTECHAMBER, with RESP HF/6-31G(d) charges calculated with Gaussian16[66] and the General Amber Force Field (GAFF).[67] Counter-ions (Na^+) were added to neutralize the overall charge and the complete systems were placed in with TIP3P water boxes with a minimum distance of 12 Å between the protein-surface and the side of the box.

Four consecutive minimizations stages were performed to remove clashes prior to the MD simulation. In these four stages the minimization procedure was applied to the following atoms of the system: 1-water molecules (2500 steps); 2-hydrogens atoms 2500 steps); 3-side chains of all the amino acid residues 2500 steps); 4-full system 10.000 steps). The minimized systems were then subject to a molecular dynamics equilibration procedure, which was divided into two stages: in the first stage (50 ps), the systems were gradually heated to 298 K using a Langevin thermostat at constant volume (NVT ensemble); in the second stage (50 ps) the density of the systems was further equilibrated at 298 K. Finally, molecular dynamic production runs were performed for 100 ns. These were performed with an NPT ensemble at constant temperature (298 K, Langevin thermostat) and pressure (1 bar, Berendsen barostat), with periodic boundary conditions, with an integration time of 2.0 fs using the SHAKE algorithm to constrain all covalent bonds involving hydrogen atoms. A 10-Å cutoff for nonbonded interactions was used during the entire molecular simulation procedure. Final trajectories were analyzed in terms of Root-Mean-Square Deviation (RMSD), confirming that both systems were well equilibrated after the initial 10 ns. The last 90 ns of the simulation were considered for hydrogen bonding analysis, and cluster analysis of the conformations generated. This overall procedure has been previously used with success in the treatment of several biomolecular systems.[68-74]

In order to estimate the binding free energies of the ligands **6** and **8** toward acetylcholinesterase and odorant binding protein 1, the Molecular Mechanics - Generalized Born Surface Area (MM-GBSA) method[75] was applied, taking into account a salt concentration of 0.100 mol dm^{-3}. Also, the contribution of the amino acid residues was estimated using the energy decomposition method. From each MD trajectory, a total of 1400 conformations taken from the last 70 ns of simulation were considered for each MM-GBSA calculation.

Statistical analysis

For biological assays, the Shapiro-Wilks normality test was performed in the data to ensure that it followed a normal distribution. Comparison between the means of controls and each experimental condition was performed using ANOVA. Outliers were identified by the Grubbs' test. Data was expressed as the mean ± standard deviation (SD) of at least 3 independent experiments. GraphPad Prism 7.0 software was used, and values were considered statistically significant with a $p < 0.05$.

Conclusions

Various derivatives and analogues of carvacrol and thymol isomers were synthesized in which the hydroxyl group was transformed into ethers with different aliphatic chains, such as a propyl without or with a hydroxyl group or chlorine atom as terminals, as well as bicyclic ethers with an unsaturated chain having a carboxylic acid group. In addition, an analogue of the natural isomers was also prepared possessing all the methyl groups changed to methoxyl, 1-hydroxyethyl and (3-chlorobenzoyl)oxy groups.

All the synthesized compounds were subjected to tests of biological activity in *Sf9*, in comparison with the corresponding precursors, in order to evaluate their application as potential natural based insecticides. We identified structural changes that result in diminished insecticide activity and also highlight two molecules,

compounds **6** and **8**, that are mildly active, the latter being less toxic to human cells than the naturally-derived starting materials.

Structural-based inverted virtual screening studies with five independent methods suggest that the two most potent molecules display their insecticide activity most likely by targeting the insect odorant binding proteins and/or acetylcholinesterase. Molecular dynamics simulations and free energy calculations confirm that these two molecules bind strongly to both targets forming very stable complexes, with well defined-molecular interactions that are maintained through time.

Liposomes of egg phosphatidylcholine/cholesterol (7:3) prepared by the thin film hydration method allow high encapsulation efficiencies and a delayed release of the most active compounds, while the ones prepared ethanolic injection provide a release higher than 60% in two days.

Conflicts of interest

There are no conflicts to declare.

Author Contributions

Conceptualization, M.S.T.G., D.M.P., S.F.S. and E.M.S.C.; methodology, M.S.T.G., D.M.P., S.F.S. and E.M.S.C.; formal analysis, M.S.T.G., D.M.P., S.F.S., E.M.S.C. and R.B.P.; investigation, C.M.N., N.F.S.P., R.B.P., T.F.V. and A.R.O.R.; supervision: M.S.T.G., D.M.P., S.F.S. and E.M.S.C.; writing-original draft preparation, M.S.T.G., D.M.P., S.F.S. and E.M.S.C.; writing-review and editing, M.S.T.G., D.M.P., S.F.S. and E.M.S.C.; project administration, M.S.T.G. All authors have read and agreed to the published version of the manuscript.

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